The amount of calcium ion present in milk can be determined by adding oxalate ion, $C_2O_4^{2-}$ (in the form of its sodium salt $Na_2C_2O_4$). The insoluble compound calcium oxalate, $CaC_2O_4$, is precipitated.

$$Ca^{2+}(aq) + Na_2C_2O_4(aq) \rightarrow CaC_2O_4(s) + 2\ Na^+(aq)$$

Suppose you have a 75.0-g sample of milk and isolate 0.288 g of $CaC_2O_4$ from it. What is the mass percentage of calcium in the milk?